

Research Article

Photocatalytic Active Coatings for Lignin Degradation in a Continuous Packed Bed Reactor

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The synthesis of immobilized catalyst on porous glass support material via the sol-gel route is reported. TiO₂-P25-SiO₂ + Pt, TiO₂-P25-SiO₂, TiOSO₄·30.6 wt%, and ZnO + TiO₂-P25-SiO₂ catalysts were synthesized and a comparative study is done regarding morphology of coatings, degradation rates, reaction rates, dissolved carbon (DC), formation of peaks, and fluorescence of products formed from the photocatalytic degradation of lignin sulfonate obtained from a local paper plant. Through simultaneous reaction-extraction pathways applying dialysis filtration and highly porous polystyrene divinylbenzene adsorbent resin (HR-P) for solid phase extraction (SPE), an attempt has been made to isolate smaller molecules produced from photocatalytic degradation. Moreover relatively high lignin sulfonate (0.5 g/L) concentrations are used in the reactions. UV-Vis spectroscopy revealed a faster reduction in the concentration values for the aliphatic moiety compared to the aromatic moiety. Peaks were observed by both fluorescence spectroscopy and HPLC suggesting the production of new substances and fluorophores.

1. Introduction

There is growing awareness of the rapid depletion of fossil materials which makes it imperative to think of the development of commercial viable “green products” based on alternatives for fossil fuels. Commercial lignin is byproduct of the pulp and paper industries, wood hydrolysis industries, and pretreatment technologies to obtain cellulosic ethanol and biodiesel. More than 70 million tons of various types of lignin are produced as waste material by the paper industry yearly [1]. Coupled to that, lignin and its degradation products cannot be completely decomposed by sludge processes and are the source of chemical oxygen demand (COD) of waste water from pulp and paper mills [2]. Lignin is primarily a structural material to add strength and rigidity to cell walls and constitutes between 15 wt% and 40 wt% of the dry matter of woody plants. After cellulose, it is the most abundant

renewable carbon source on earth. Lignin can be available through various sources and biomass transformation technologies [3].

Photocatalysis, belonging to the so-called advanced oxidation processes (AOPs) [4], is a potential new transformation technology for lignin to value added products, for example, phenol, benzene, toluene, and xylene [5]. One of such a transforming method is the heterogeneous procedure in which a photocatalytically active layer is immobilized on a support material. In such a system both adsorption and photocatalysis take place and major advantages such as the reuse of catalyst; minimization of catalyst leaching or resistance to extreme physical and chemical conditions is exploited. Most publications treating lignin photocatalysis describe suspension systems [6] with the disadvantage that catalyst particles have to be recovered after reaction through downstream processes such as ultrafiltration which bring

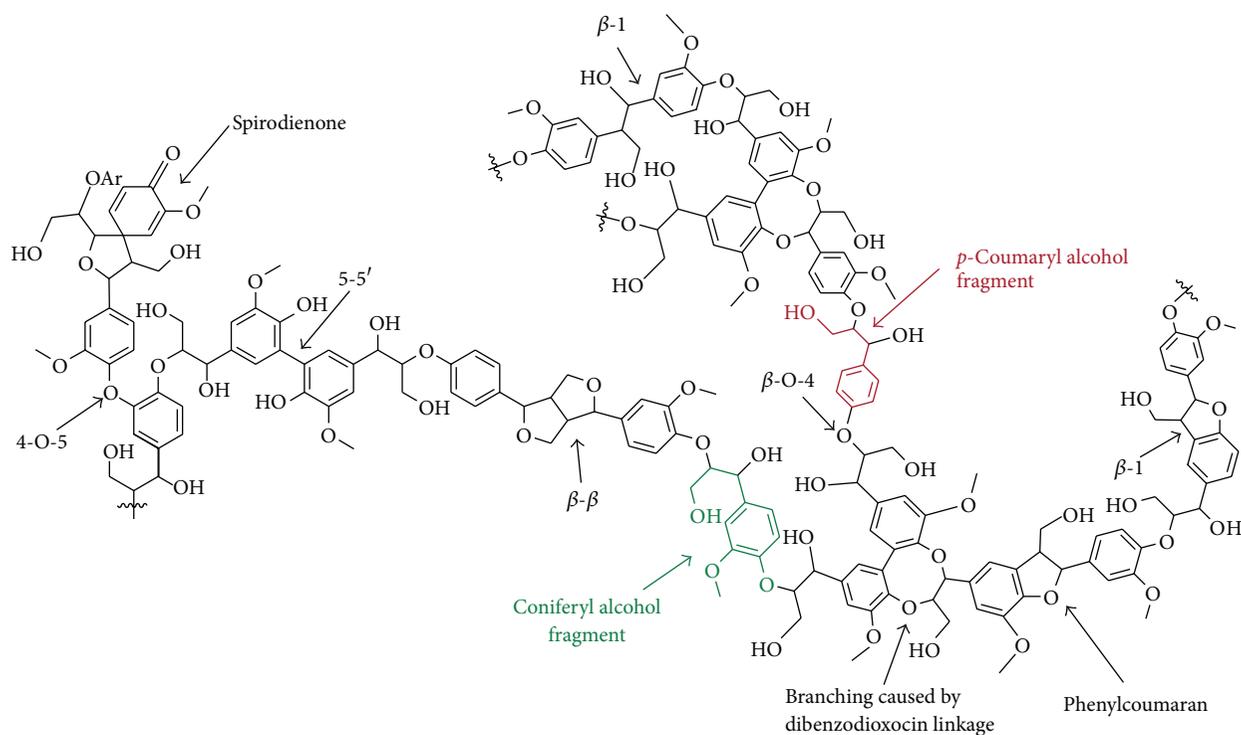


FIGURE 1: Structure of a softwood lignin molecule showing the prominent linkage types. With permission from Zakzeski et al. [40].

along additional process steps and higher cost. Besides photocatalysis, lignin can be degraded by other ways such as biochemical degradation through microbial/enzymatic attack [7–9], thermal degradation [10, 11], and electrochemical oxidation [12].

For a good catalytic efficiency in an immobilized catalytic system, a good hydrodynamic design of the system and modifications involving the utilization of the visible light spectrum by catalyst doping can be of great importance. The selection and design of the supporting material influence both the fluidic design and the available catalyst surface. The photocatalytic efficiency increases with an increase in active surface. This means the carrier material has to provide a high surface to volume ratio. Also, the materials used as supports for photocatalytic purposes must also fulfill stringent optical properties such as transparency to UV-A light, in order to achieve high photonic efficiencies and to minimize the energy consumption required for the illumination.

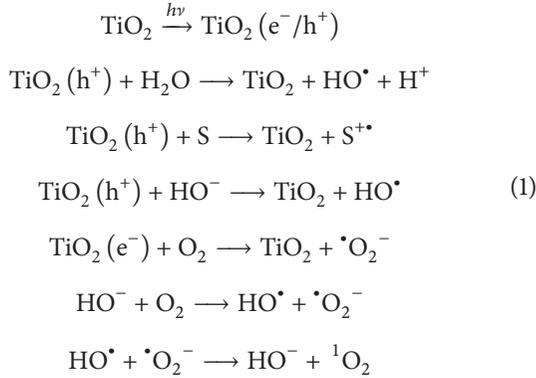
The goal of this work is to describe stable catalytic coatings produced on porous glass material and report their photocatalytic activity which was studied by degradation experiments of lignin under UV light illumination. Sintered glass particles were coated to exploit the advantages of a heterogeneous system and a good fluidic design was attained by the porosity of the glass particles. Moreover, parameters such as flow characterization and irradiance which are required for the rigorous design and the scaling up of a photocatalytic reactor in order to accomplish a commercial application are treated in this paper. Additionally a simultaneous reaction-extraction pathway, applying dialysis filtration and highly

porous polystyrene divinylbenzene adsorbent resin (HR-P), was applied. In this work, a relatively high concentration of lignin sulfonate (0.5 g/L) obtained from a local paper company is used.

Furthermore, this paper presents a reaction design which can remediate the pollution problem in waste water discharged from pulp and paper mills on the one hand and, on the other hand, gives room to further exploit the degradation of lignin as an alternative chemical source. Results obtained from the time dependent degradation gives an opportunity to achieve defined chemical structures when isolated. Peaks observed by both high performance liquid (HPLC) and fluorescence chromatograms suggest the production of new substances and lignin sulfonate.

1.1. Degradation and Mineralization of Lignin. Lignin degradation is generally in the range of lower energy (between 300 and 400 nm) because of its multifunctional character [13–15]. For a better illustration, Figure 1 shows the structure of a softwood lignin fragment containing all prominent linkage types. During the photodegradation process of lignin, radical sites and excited α -carbonyl are involved. Excited α -carbonyl groups are able to abstract hydrogen from some lignin units, preferably phenolic hydrogen, initiating an oxidative chain reaction with the participation of ground-state oxygen, leading to the fragmentation and the formation of new chromophores, dimers, and sometimes oligomers as major products [16]. The action of hydroxyl radicals can induce the formation of radical sites on a substrate, thus inducing further oxidation or disproportionation reactions.

The process can be accelerated by oxygen [17]. Machado et al. [18] report a rapid initial decrease of phenolic hydroxyl groups followed by a clear tendency to increase explained as an effect of incorporation of hydroxyl groups on aromatic rings, occurring by hydroxylation and/or demethoxylation [17]. The below equation summarizes the formation of singlet oxygen, hydroxyl, and superoxide radicals as principal reactive species in a photocatalytic process [19, 20]. S stands for the lignin substrate while $\text{TiO}_2(\text{h}^+)$ and $\text{TiO}_2(\text{e}^-)$ represent the electron-deficient and electron-rich parts in the structure of TiO_2 , respectively. Consider



2. Experimental

2.1. Flow Measurements. The reaction design consisted of TiO_2 coatings on sintered glass (150–250 μm nominal pore size), packed in a borosilicate tube 26 cm long, 1 cm outer diameter, and 6 mm internal diameter. Glass particle grain size was between 200 and 400 μm . The packed bed reactor design had advantages such as a constant plug flow and therefore a narrow residence time distribution and good mixing condition with the sintered glass acting as static mixer.

For fluid flow through a bed of approximately spherical particles, the Reynold number (Re) used to characterize laminar or turbulent flow regimes is defined as follow [21]:

$$\text{Re} = \frac{\rho V D}{\mu (1 - \varepsilon)}, \quad (2)$$

where ρ is the density of lignin sulfonate solution (kg/m^3), μ is the dynamic viscosity of 0.5 g/L lignin sulfonate solution (Pa·S), ε is the voidage and V is the superficial velocity (m/s). The voidage was calculated from the correlation equation of packed beds [22].

The voidage was calculated from the correlation equation of packed beds [22]:

$$\varepsilon = 0.293 + (0.684) \text{dr}^{-0.85} \cdot \frac{1}{(1.837\text{dr} - 1)^{0.5}}, \quad (3)$$

where $\text{dr} = \text{dt}/\text{dp}$, dt is diameter of tube and dp is diameter of particle.

The particle must be spherical and $\text{dr} > 1.86$ [22].

The superficial velocity V was calculated from

$$V = \frac{\text{volume flow rate}}{\text{cross sectional area of tube}}. \quad (4)$$

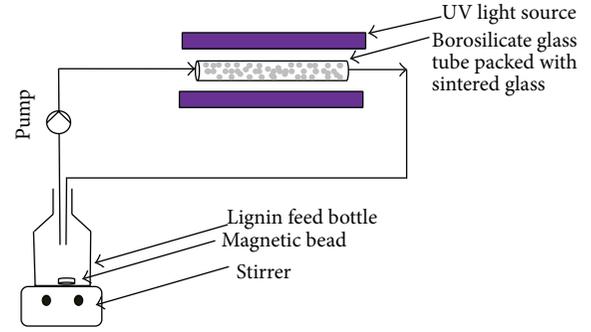


FIGURE 2: Experimental setup for the photocatalytic degradation of lignin.

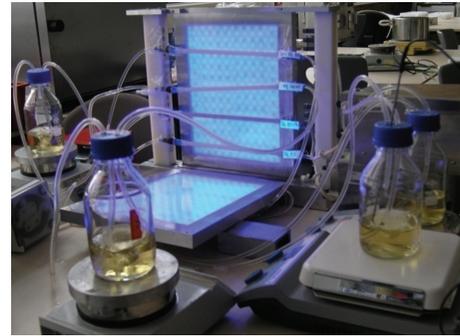


FIGURE 3: Photocatalytic reactor.

2.2. Photocatalytic Experiments. The degradation experiments of lignin sulfonate were carried out with a starting concentration of 0.5 g/L in 200 mL deionized water. The tube was placed between two planar dielectric barrier discharge lamps. The broad surface of the light source insured incident light is well distributed all over the surface of the glass tubes. Figure 2 depicts the reaction setup while Figure 3 illustrates the reactor used for the experiments.

The reaction procedure involved 3 pathways. In the first one, lignin sulfonate solution was pumped through the reactor using roller pumps in a continuous manner with a flow rate of 22.5 mL/min. Samples were collected at intervals up to a period of 20 h reaction time. Reactions were carried out at room temperature. Blank experiments were performed in the absence of UV light illumination as well as with uncoated sintered glass tubes. For all experiments, lignin was recirculated through the reactor for 60 min before UV-light was put on. This was to ensure uniform wettability within the packed bed and same starting reaction conditions for all experiments. The second procedure involved a reaction-extraction pathway in which extraction followed after reaction with the use of highly porous polystyrene divinylbenzene adsorbent resin (HR-P). In the third procedure, the reactor was connected to a dialysis filter fitted with a HRP extraction column operating in a cycle process. This is depicted in Figure 4. Its aim was to directly extract smaller molecules produced and preferably aromatic and phenol-like compounds from the aqueous medium.

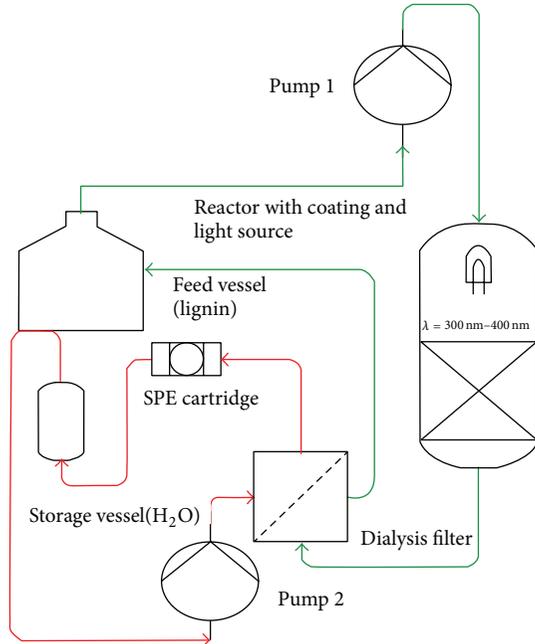


FIGURE 4: Photocatalytic cycle process with integrated solid phase extraction and dialysis membrane.

2.3. Analytical Method. The concentration of lignin in the reaction mixture was determined by correlating absorbance of an aliquot solution to the calibration curve of lignin at different concentrations. This was done using a UV-visible (Thermo-Scientific Genesys 10S) spectrophotometer (at 203 nm and 280 nm) with deionized water as reference. The percentage degradation was calculated as follows:

$$\text{Degradation (\%)} = \left(\frac{C - C_0}{C_0} \right) \times 100. \quad (5)$$

C_0 is the initial concentration of lignin and C is the concentration of lignin after photo-irradiation. The photocatalytic decomposition of the organic molecules follows the Langmuir-Hinshelwood kinetics, which may be represented as [23, 24]

$$\frac{-dC}{dt} = -kCC_{HO\cdot} \quad (6)$$

where C is the concentration of lignin and $C_{HO\cdot}$ the hydroxyl radical concentration. By the pseudostationary hypothesis (i.e., the $C_{HO\cdot}$ can be considered to be constant because of the continuous hydroxyl ion generation by photocatalysis) the rate depends only on the lignin concentration [25]. The rate expression can be written as follows:

$$-\ln \left(\frac{C}{C_0} \right) = kt, \quad (7)$$

where the linear plot of $\ln(C/C_0)$ versus t gives the rate constant, k (slope).

Dissolved carbon (DC) content of lignin sulfonate over time was done using a TOC analyser (Shimadzu TOC-5000A). High performance liquid chromatography (HPLC)

and fluorescence measurements were done on an HP1090 liquid chromatograph and HP1040A programmable fluorescence detector, respectively.

2.4. Material. Lignin sulfonate waste water was obtained from a local producing paper company (Sappi-Ehingen AG). Cross flow filtration was applied using ceramic membranes of 0.2 μm and 20 kD membrane to isolate lignin. The concentrated lignin sulfonate solution was then dried in a laboratory drying chamber at 40°C for several days in order to avoid mold and to ensure constant concentration during dissolution. ZnO (5 m^2/g) and chloroplatinic acid hydrate (99.999%) were purchased from Sigma-Aldrich, Germany, and titanium (IV) oxide sulfate dihydrate (titanyl oxy sulfate) was purchased from Alfa Aesar. Aeroxide TiO₂-P25 (surface area 50 m^2/g and average particle size 30 nm) was obtained from Evonik and consisted of 70% anatase and 30% rutile. tetraethyl orthosilicate (TEOS) was purchased from Merck. Borosilicate glass tubes containing sintered glass particles were purchased from ROBU, Germany. Polystyrene divinylbenzene adsorbent resin, Chromabond (HR-P) was purchased from Macherey Nagel. The dialysis filter was an ultraflux AV Paed (surface area 0.2 m^2/g and with high cut off, average pore size < 0.02 μm). UV-A meter was purchased from Dr. Honle UV Technology. All chemicals were used as received.

2.5. Preparation of Catalyst and Coating Procedure. The following catalysts were synthesized: TiO₂-P25-SiO₂ + Pt, TiO₂-P25-SiO₂, ZnO + TiO₂-P25-SiO₂, and TiOSO₄ 30.6 wt%. A detailed description of the catalysts synthesis and coating has been described elsewhere [26].

3. Results and Discussion

3.1. Morphology of Coating and Surface Area. Scanning electron microscopy (SEM) was the primary tool for characterizing the surface morphology and fundamental physical properties of the material surface such as porosity and distribution of the coating. In Figure 5, SEM images are shown for uncoated and coated glass with different catalysts. Samples consisted of sintered glass plates cut to a size of 1 cm^2 dipped into the gel or suspension. Drying and calcination were done same as was the case with borosilicate glass tubes [26]. The micrographs for the uncoated sintered glass in Figures 5(a) and 5(b) show structures formed by glass particles of approximately 200–400 μm in size sintered together. Figures 5(c)–5(h) shows micrographs of the coated glass material. Figures 5(e) and 5(h) consist of a double coating whereby TEOS was first coated and TiO₂-P25 follows. Significant agglomeration of the particles is also observed. Some fissures are also seen at joints between sintered particles and on the glass particles as a result of drying. It is worthwhile noting that these micrographs originate from sintered glass plates whereby the drying mechanism could be harsher as heat radiations directly strike on them as compared to a borosilicate glass tube encasing sintered glass particles.

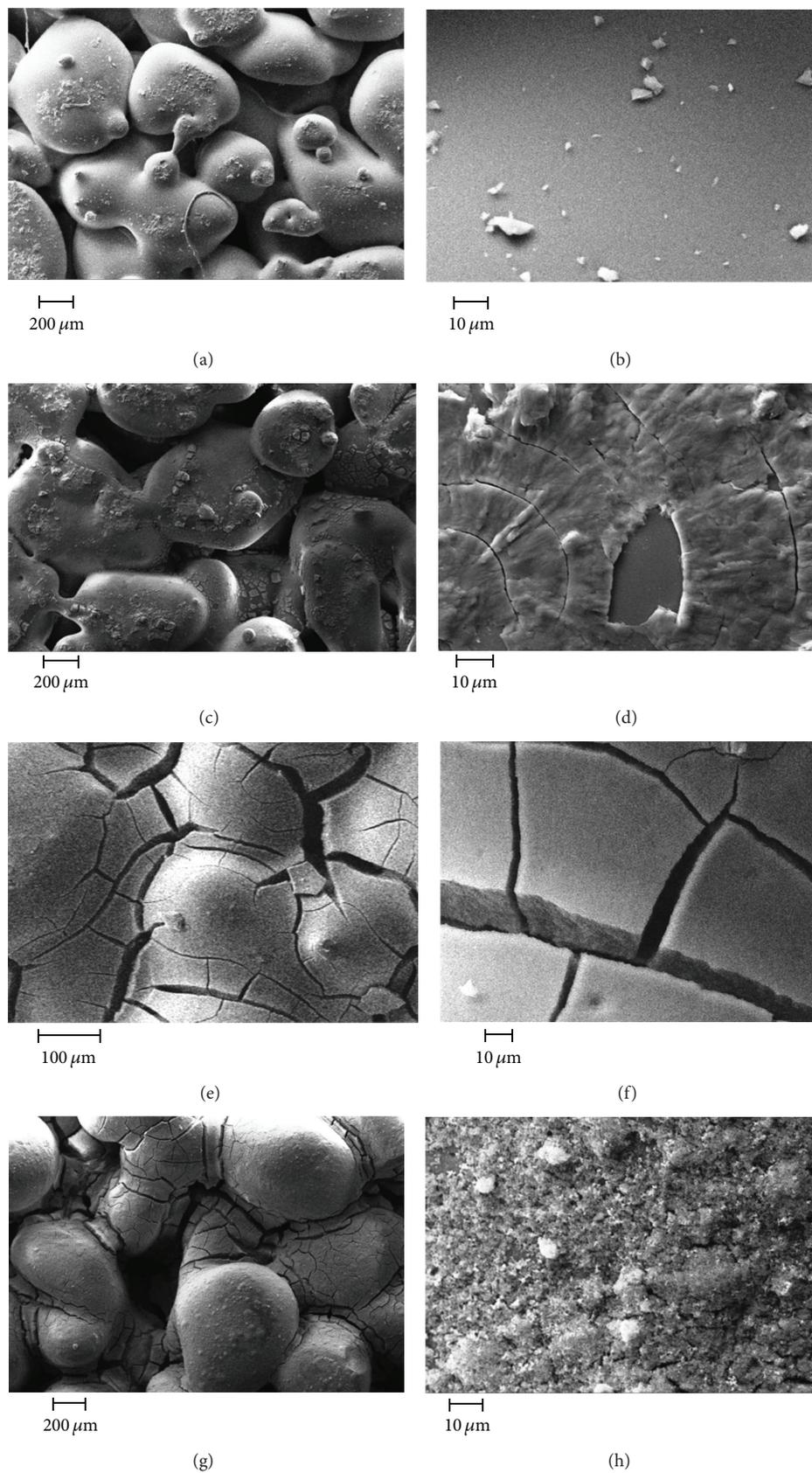


FIGURE 5: SEM images for uncoated borosilicate sintered glass ((a) and (b)), TiO₂(TiOSO₄ 30.6 wt%) coating on borosilicate glass ((c) and (d)), TiO₂-P25-SiO₂ ((e) and (f)), and ZnO + TiO₂-P25-SiO₂ ((g) and (h)).

TABLE 1: Parameters used to calculate the flow characteristic of photocatalytic reactor.

$R_{i,tube}$	$A_{tube} (m^2)$	$\rho_{lignin} (kg/m^3)$	$V_{lignin} (m/s)$	$D_{particles} (\mu m)$	$\mu_{lignin} (kg/ms)$	ϵ (voidage)	Re
0.30 cm	2.83×10^{-5}	1000	0.013	300	1×10^{-4}	0.30	56

It can be observed that the coatings have a crystalline structure and are considerably well distributed all over the surface of the sintered glass. Rutile is known to be the stable phase of TiO_2 whereas anatase and brookite are metastable at all temperatures and transform to rutile when they are heated. From the calcination temperature of $500^\circ C$, the crystal phases involved both rutile and anatase since the transformation of anatase to rutile is in the temperature range between 600 and $700^\circ C$ [27]. Moreover anatase is typically the majority product of inorganic syntheses and is the main constituent of nanocrystalline materials [28]. Aeroxide TiO_2 -P25 obtained from Evonik consisted of 70% anatase and 30% rutile.

3.2. Reactor Characterization. For packed bed reactors, laminar conditions apply when $Re < 2000$ and fully turbulent from 2000 [21]. Using (2), a value of 56 was obtained which characterizes a laminar flow. Table 1 outlines the flow characterization values.

The weight of coating on the glass tubes was in the range of 0.12 g–0.19 g. The BET specific surface area of a borosilicate tube containing the sintered glass particle was $0.015 m^2/g$ making a total of $0.17 m^2$ for 11.5 g glass particles in the borosilicate glass tube.

UV-light was from two planar dielectric barrier discharge lamps (Osram Planon) emitting UV-light in the range of 280–420 nm. The irradiance determined by UV-A radiometric measurements was between 30 and $40 W/m^2$.

4. Results and Discussion

4.1. Degradation Experiments. Figure 6 shows a time dependent UV-Vis spectra of lignin sulfonate. Absorption peaks were observed at 203 nm and 280 nm while a shoulder at around 230 nm was observed. Lignin absorbs UV light with high molar extinction coefficients because of the several methoxylated phenylpropane units of which they are composed [29]. The absorption peaks decrease gradually indicating the decomposition of lignin sulfonate and hence deterioration of the present chromophore groups. Peaks around 203 nm correspond to portions of unsaturated chains [29, 30] while those around 280 nm correspond to unconjugated phenolic hydroxyl groups [31] and aromatic rings [30] of lignin sulfonate. Ohnishi et al. [30] report the absorption tailing arising from the color of lignin.

Preliminary reactions carried out involved lignin sulfonate degradation experiments performed on uncoated glass in the presence of UV-light as well as on coated glass particles in the absence of UV-light. Results obtained revealed no significant change in absorbance over a 20 h period indicating no major degradation and adsorption. Houas et al. [24] note that, at catalytic surfaces experiencing unfavourable conditions such as repellent interactions, there may be no

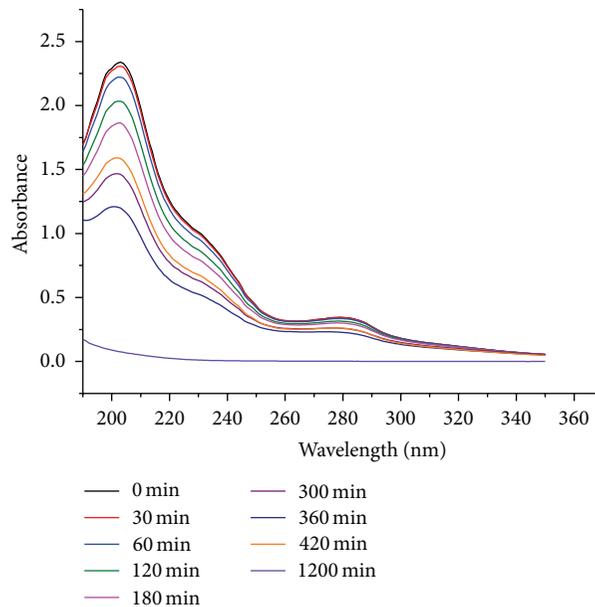


FIGURE 6: Time dependent UV-Vis absorption spectra of aqueous lignin sulfonate solution. Concentration: 0.5 g/L in 200 mL, reaction time: 20 h, catalyst: sol-gel derived TiO_2 -P25- SiO_2 , reaction at room temperature, and UV radiation (280–420 nm).

adsorption. It can further be argued that because of the nature of lignin used in this work (lignin sulfonate) the sulfonate groups attached to lignin render it anionic. Likewise because of the basic nature of the reaction medium (pH 9) and the negatively charged hydroxyl (OH^-) ions generated from photochemical reactions, adsorption conditions do not prevail.

Figures 7 and 8 depict the degradation of lignin sulfonate by different catalysts at wavelengths 203 nm and 280 nm, respectively. Degradation of lignin sulfonate prevailed in the presence of a catalyst. An abrupt degradation increase was observed for $TiOSO_4$ -30.6 wt% and $ZnO + TiO_2$ -P25- SiO_2 catalysts between 240 min and 300 min. Here, the slow reaction at the beginning is the rate determining step [32]. In this step, processes such as absorption of photons of light by catalyst and production of electrons and other oxidizing species take place before a faster reaction takes place [33]. An almost linear degradation pattern was noted for TiO_2 -P25- $SiO_2 + Pt$ and TiO_2 -P25- SiO_2 throughout the complete reaction time. Degradation rates of 97%, 95%, 87%, and 51% were attained by coatings TiO_2 -P25- $SiO_2 + Pt$, TiO_2 -P25- SiO_2 , $ZnO + TiO_2$ -P25- SiO_2 , and $TiOSO_4$ -30.6 wt%, respectively, after 20 h reaction time for the straight chain lignin sulfonate moiety (Figure 7). Similar degradation values are

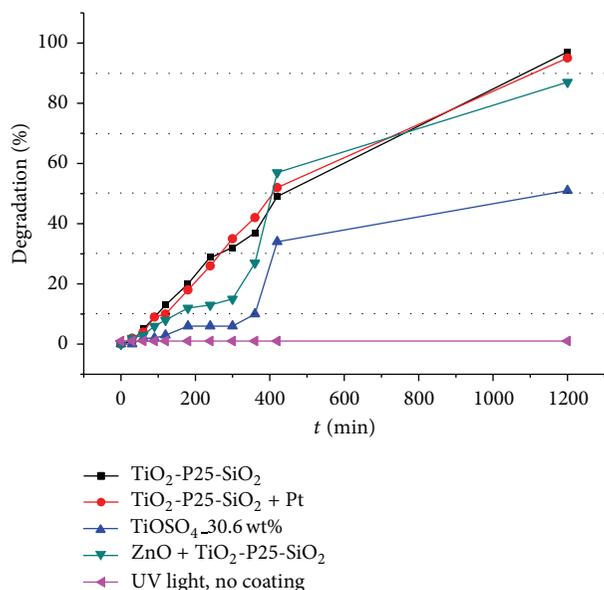


FIGURE 7: Effect of different catalyst on percentage degradation of lignin sulfonate under UV light at wavelength 203 nm (lignin sulfonate concentration: 0.5 g/L in 200 mL, reaction time: 20 h, reaction at room temperature, and UV radiation (280–420 nm)).

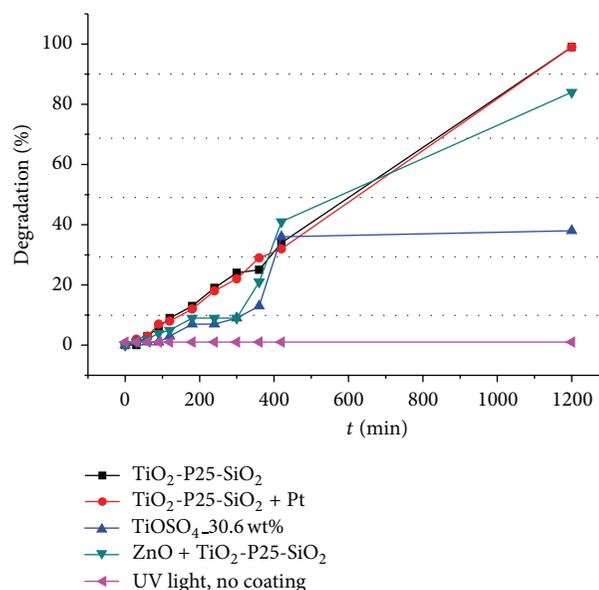


FIGURE 8: Effect of different catalyst on percentage degradation of lignin sulfonate under UV light at wavelength 280 nm (lignin sulfonate concentration: 0.5 g/L in 200 mL, reaction time: 20 h, reaction at room temperature, and UV radiation (280–420 nm)).

also observed for the aromatic moiety after 20 h (Figure 8). TiOSO₄_30.6 wt% catalyst was less reactive compared to the others because of the relatively low degradation values it produced. Less than 40% degradation was achieved after 420 min for both the aromatic and aliphatic moiety. Comparing the degradation rates of the aromatic moiety against that of the straight chain moiety after 420 mins, it was observed that higher degradation values were attained by the aliphatic moiety. In detail, for the aliphatic moiety degradation rates of 49%, 52%, 57%, and 34% were noted while for the aromatic moiety degradation rates of 34%, 32%, 41%, and 36% were observed for TiO₂-P25-SiO₂ + Pt, TiO₂-P25-SiO₂, ZnO + TiO₂-P25-SiO₂, and TiOSO₄_30.6 wt%, respectively. This suggests a faster transformation of the aliphatic side chains as compared to the aromatic species. A similar conclusion was arrived at by Tanaka et al. [34] whereby FTIR measurements were applied.

Comparing the degradation values of lignin sulfonate by using plain TiO₂-P25-SiO₂ and that of the combination between ZnO and TiO₂-P25-SiO₂, a higher degradation rate was observed for ZnO and TiO₂-P25-SiO₂ after 420 mins, both for the straight chain moiety and aromatic moiety. However after 20 h reaction time, TiO₂-P25-SiO₂ had a higher degradation value, both for the straight chain moiety and aromatic moiety. Degradation rates of 49% and 57% were observed after 420 minutes for TiO₂-P25-SiO₂ and ZnO + TiO₂-P25-SiO₂ catalysts, respectively. Meanwhile after 20 h reaction time, values of 97% and 87% are attained by TiO₂-P25-SiO₂ and ZnO + TiO₂-P25-SiO₂ catalysts, respectively. This means ZnO + TiO₂-P25-SiO₂ catalyst was more active during the first 420 mins and its activity decreased in the proceeding time.

Adding (doping) Pt as metal ion to TiO₂-P25-SiO₂ slightly improved the degradation values compared to that of plain TiO₂-P25-SiO₂ catalyst. Values of 49% and 52% are attained after 420 minutes reaction time for TiO₂-P25-SiO₂ and TiO₂-P25-SiO₂ + Pt catalysts, respectively, whereas, after 20 h reaction time, values of 97% and 95% are attained by TiO₂-P25-SiO₂ and TiO₂-P25-SiO₂ + Pt catalysts, respectively. These results indicate Pt contributed little or no change in degradation rate which is contradictory to some published results on the effect of doping [35, 36]. However, for reactions in which doping enhances degradation, a threshold dosage has to be considered [36].

In order to effectively quantify lignin degradation, dissolved carbon (DC) measurements were done (Figure 9). DC describes the amount of carbon bound in a compound. DC in lignin sulfonate of initial concentration 500 mg/L decreased from 185 ppm to 144 ppm, 121 ppm, 159 ppm, and 143 ppm after 420 mins for the catalysts TiO₂-P25-SiO₂, TiO₂-P25-SiO₂ + Pt, ZnO + TiO₂-P25-SiO₂, and TiOSO₄_30.6 wt%, respectively. This value further decreased to 31 ppm, 33 ppm, 122 ppm, and 71 ppm making a DC removal of 84%, 82%, 61%, and 34% for TiO₂-P25-SiO₂, TiO₂-P25-SiO₂ + Pt, ZnO + TiO₂-P25-SiO₂, and TiOSO₄_30.6 wt%, respectively. Hence, TC removal was in the order of TiO₂-P25-SiO₂ + Pt ~ TiO₂-P25-SiO₂ > ZnO + TiO₂-P25-SiO₂ > TiOSO₄_30.6 wt%. Decrease in DC has been reported to generate carbon dioxide and a small amount of carbon monoxide as the main gaseous products [35].

The rate constant values for the different catalysts are shown in Figure 10. Reaction rate was in the order of TiO₂-P25-SiO₂ + Pt ~ TiO₂-P25-SiO₂ > ZnO + TiO₂-P25-SiO₂ >

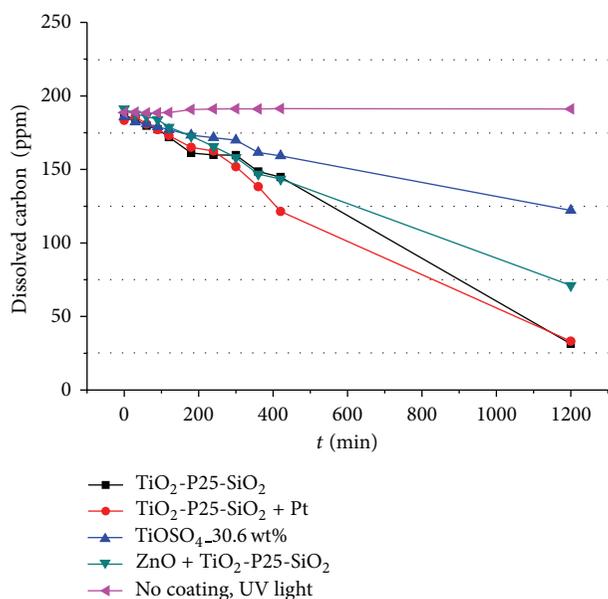


FIGURE 9: Removal of dissolved carbon in lignin sulfonate. Concentration: 0.5 g/L in 200 mL, reaction time: 20 h, reaction at room temperature, and UV radiation (280–420 nm).

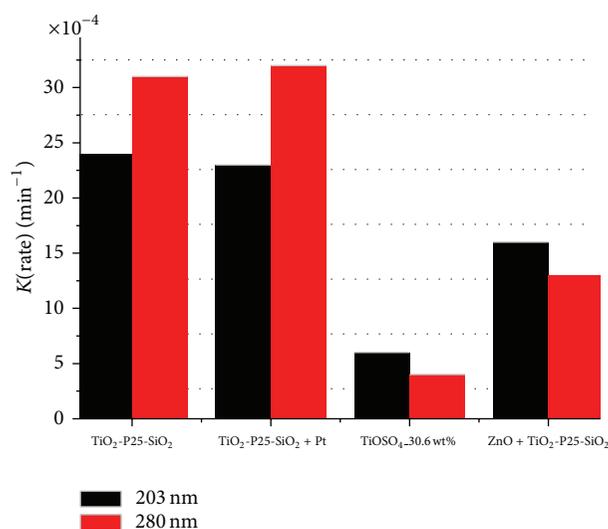


FIGURE 10: Reaction rates of the photocatalytic degradation of lignin sulfonate (500 mg/L) under UV light on immobilized catalysts. Concentration: 0.5 g/L in 200 mL, reaction time: 20 h, reaction at room temperature, and UV radiation (280–420 nm).

TiOSO₄_30.6 wt% for both the aromatic and aliphatic side chain wavelength regions.

Comparing ZnO + TiO₂-P25-SiO₂ and TiO₂-P25-SiO₂ and adding ZnO to TiO₂-P25-SiO₂ had an adverse effect on reaction rate probably because ZnO occupies some of the active sites on the catalyst support and it does not optimally absorb UV light in the domain of the wave lengths transmitted by the photo reactor (280 < λ < 420 nm) and, as a consequence, limitations in electron transfer between the band gaps [37].

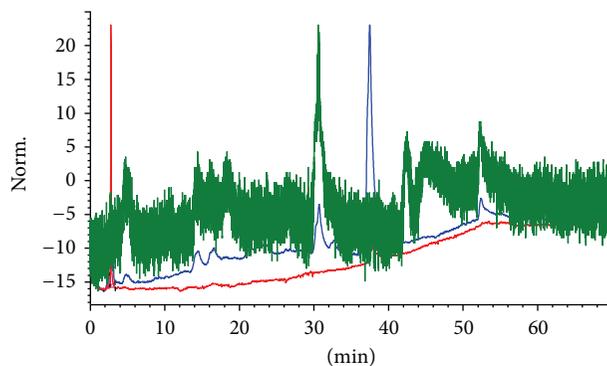


FIGURE 11: HPLC chromatogram and fluorescence peaks of lignin sulfonate solution. Concentration: 200 mL of 0.5 g/L, reaction time: 4 h, catalyst: TiO₂-P25-SiO₂, reaction at room temperature, presence of UV radiation, simultaneous reaction, dialysis filter, and SPE with HP-R cartridge during reaction.

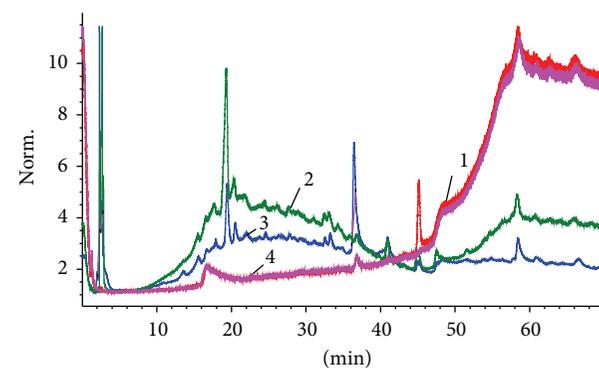


FIGURE 12: Fluorescence chromatogram of lignin sulfonate degradation product. Emission spectrum, λ_{em} : 240 nm, λ_{ex} : 330 nm, concentration: 100 mL 0.5 g/L, reaction time: 20 h, catalyst: TiOSO₄_30.6 wt%, and reaction at room temperature. (1) Absence of UV radiation and SPE with HP-R cartridge after reaction suspended in methanol. (2) Presence of UV radiation and SPE with HP-R cartridge after reaction suspended in methanol. (3) Presence of UV radiation and simultaneous SPE with HP-R cartridge during reaction. (4) Presence of UV radiation and aqueous raffinate from continuous extraction.

4.2. Fluorescence and HPLC Results. Fluorescence spectroscopy was used as a means to identify nonaliphatic component in the complex mixture of lignin degradation products. Fluorescence emission in lignin is attributed to aromatic structures such as conjugated carbonyl, biphenyl, phenylcoumarone, and stilbene groups [38, 39]. New fluorophores were detected in the time ranging from 10 min to around 40 min (Figure 12) with application of TiOSO₄_30.6 wt% catalyst. With the use of TiO₂-P25-SiO₂ catalyst, new fluorophores were detected as soon as 5 min to around 55 mins (Figure 11).

Comparing pathway (2) and pathway (3) from Figure 12, it was observed that the fluorophores produced had similar retention times and similar peak areas. This indicates that it did not matter if extraction was done after the reaction or

during the reaction. No peaks were observed for (1) reinforcing the necessity of UV light for lignin sulfonate degradation. The absence of peaks for (4) suggests that degradation was complete and all produced substances were adsorbed on the SPE cartridge. Moreover, tailing was observed for samples (1) and (4) that did not contain degradation products that dissolved in it.

Figure 11 depicts an HPLC and fluorescence chromatogram in which both measuring apparatus were coupled. The peak at around 2 min retention time arises from methanol used as eluting agent during SPE. Peaks were observed on both chromatograms suggesting the production of new substances in which there are some generated fluorescence peaks. However some fluorescence peaks appear only on the fluorescence chromatogram. Also broad peaks were found both on the HPLC and fluorescence chromatograms. During measurements (diode array detector) DAD wavelength was set at 240 nm and not all degradation products could be detected at this particular wavelength. In fact, one of the challenges to analyze lignin through HPLC is to set the DAD wavelength for maximum detection. An analogous scenario was observed during fluorescence in which emission wavelength was set at 240 nm and excitation wavelength set at 330 nm. These difficulties arise because of the vast range of possible products arising from lignin degradation. Because of that, distinct emission decay peaks cannot be produced due to superimposition of the different fluorophores.

5. Conclusion

Commercial obtainable TiO_2 -P25 widely applied in suspension systems was effectively immobilized on sintered glass support with the aid of tetraethyl orthosilicate (TEOS) solution which acted as support material. The catalytic films synthesized had very good adherence on the sintered glass and the films retained a constant efficiency when reused after being washed with water. The coatings were considerably well distributed all over the surface of the sintered glass. Moreover, the reactor design in this work effectively degraded relatively high concentration of lignin sulfonate solutions (500 mg/L) from paper waste water.

UV-Vis spectroscopy revealed a faster transformation of the aliphatic side chains as compared to the aromatic species. Adding Pt as metal ion to TiO_2 -P25- SiO_2 catalyst showed a negligible effect on degradation rates. The reaction rate of the catalytic systems was in the order of TiO_2 -P25- SiO_2 + Pt \sim TiO_2 -P25- SiO_2 > ZnO + TiO_2 -P25- SiO_2 > TiOSO_4 30.6 wt%.

Peaks were observed on HPLC and fluorescence chromatograms suggesting the production of new substances and fluorophores. Through simultaneous reaction-extraction pathways applying dialysis filtration and highly porous polystyrene divinylbenzene adsorbent resin (HR-P), an attempt has been made to isolate smaller molecules produced from photocatalytic degradation.

Conflict of Interests

The authors declare that there is no conflict of interests regarding the publication of this paper.

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